

TABLE OF SOLUBILITY OF SOME COMMON COMPOUNDS

1. • All common compounds made up of alkali metal ions (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , or Fr^+) or NH_4^+ and negative ions are *soluble*.
2. • All compounds containing $\text{C}_2\text{H}_3\text{O}_2^-$ or NO_3^- and a positive ion are *soluble*.
3. • All compounds containing SO_4^{2-} are *soluble except* when combined with ions of Ba^{2+} , Sr^{2+} , or Pb^{2+} .
4. • All compounds containing Br^- , Cl^- , or I^- are *soluble except* when combined with ions of Ag^+ , Hg_2^{2+} , or Pb^{2+} .
5. • Compounds of HCO_3^- with alkaline earth metal ions (Be^{2+} , Mg^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Ra^{2+}) and alkali metal ions (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , or Fr^+) or NH_4^+ are *soluble*. Compounds of HCO_3^- are *insoluble* with all other positive ions.
6. • Compounds with OH^- are *soluble only* with Ba^{2+} , Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Fr^+ , or NH_4^+ . Compounds with OH^- are *insoluble* with all other positive ions.
7. • All compounds with CO_3^{2-} , PO_4^{3-} , or S^{2-} are *insoluble except* with alkali metal ions (Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , or Fr^+) or NH_4^+ .